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Calorimetry Reactions Are Usually Done At Either
Constant V (in A Closed Container) Or Constant P (open
To The Atmosphere). In Either Case, We Can Measure
Q By Measuring A Change In T (assuming We Know
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Lid, Thermometer, 50- ML And 100-mL Graduated
Cylinders, Weighing Boat, 1.0 M HCl, 1. Mar 4th,
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Lhc History Question 1 Question 2 Question 3 Question 4(x) Name The Religious Order Founded By St Ignatius Loyola To Promote The Catholic Religion During The Counter-Reformation. (2) (vii) Explain Why There Was Apr 4th, 2024 CHEMISTRY 30 Assessment Enthalpy Change And Calorimetry ... A Student Uses A Simple "coffee Cup" Calorimeter To Determine The Molar Enthalpy Of Solution For Solid Ammonium Nitrate. The

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 $4.18 \text{ J } \cdot \text{C}^{-1} \cdot \text{g}^{-1}$ $Q = 8.30 \text{ C}$ $T_f = T_i + T = 21.4 + 8.30 = 29.7 \text{ C}$ 7) In Order To Write The Thermochemical Equation, You First Need To Calculate The MOLAR ENTHALPY CHANGE. 1) Calculate The Heat Change, Q, For The Surroundings - In This Case, The Surroundings Is The TOTAL VOLUME OF THE SOLUTION
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