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Can Assume The Ideal Gas Equation Of State. $V = RT/P \Rightarrow (\partial V / \partial T)_P = R/P$. Substitution $\Rightarrow \Delta S = C_p \ln(T_2/T_1) - R \ln(P_2/P_1)$ Applying $\Delta S = 0 \Rightarrow \ln(T_2/T_1) = R/C_p \ln(P_2/P_1) \Rightarrow (T_2/T_1) = (P_2/P_1)^{(C_p - C_v)/C_p}$ T T P
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