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(1) $C_{i,t+1} = C_{i,t} - W_{i,t} + f_{i,t} [C_{i,t+1} - (C_{i,t} - W_{i,t})]$, $f_{i,t} > 0$ (1) $C_{i,t+1} = C_{i,t} - W_{i,t} + f_{i,t}$
 $P_{i,t} [C_{i,t+1} - (C_{i,t} - W_{i,t})]$, $f_{i,t} > 0$, Where $C_{i,t}$ is the actual stock of plant and
equipment, $W_{i,t}$ is depreciation, and $C_{i,t+1}$ is desired plant and equipment. The
subscripts refer to firm and year. Equation (1) indicates that the stock of capital
 $W_{i,t}$ 2th, 2024 Chemistry Conversion Factors Practice Problems With
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Constants Chemistry Conversion Factors And Constants METRIC CONVERSION
FACTORS Prefix Abbreviation Conversion Factor For Example" For Example" Mega-
M 1000000 10⁶ 1 Megabyte = 1 X 10⁶ Bytes 1 Byte = 10⁻⁶ Megabytes Kilo- K 1000
10³ 1 Kilometer = 1000 Meters 1 Meter = 0.001 Kilometers Deci- D 1th,
2024 Chemistry: Conversion Factors Convert 2875 cm³ To Liters. 6. The density of
lead (Pb) is 11.34 g/cm³. Find the density of Pb in kg/dm³. 7. Convert 5.2 cm of
magnesium (Mg) ribbon to mm of Mg ribbon. 8. Convert 0.049 kg sulfur (S) to g
of S. 9. Convert 0.020 kg of tin (Sn) to mg of Sn. 10. Convert 150 mg of
acetylsalicylic acid (aspirin) to g of aspirin. 1th, 2024.

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mol 2. How many moles are in 79.85 g Fe₂O₃? Molar mass = 159.7 79.85g / 159.7 =
0.5 mol 3. How many molecules are in 720 g of C₆H₁₂O₆? Molar mass ... 3th,
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