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Relationship, So 0.0322 Mol Of Aniline Produces 0.0322 Mol Of Acetanilide. The Amount Of Acetic Anhydride Does Not Play A Role, Since It Is In Excess. Then,  $n(\text{N}) = 0.0322 \text{ mol aniline} \times \frac{1 \text{ mol Acetanilide}}{1 \text{ mol Aniline}}$ . The Molar Mass Of Acetanilide:  $(12 \times 8) + (1 \times 9) + (14 \times 1) + (16 \times 1) + 135 = 179 \text{ g/mol}$ .  
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2. Determination Of Cu(II) Content In  $\text{Cu}(\text{gly})_2 \cdot x\text{H}_2\text{O}$

(1) Weigh 1.0 1.2 G (precision Of  $\pm 0.0002$  G) Of  $\text{Cu}(\text{gly})_2 \cdot x\text{H}_2\text{O}$  With A Dry 100  $\text{cm}^3$  Beaker As The Container. (2) Dissolve It With 40  $\text{cm}^3$  Of Water And 8  $\text{cm}^3$  Of  $\text{H}_2\text{SO}_4$  (1.0  $\text{mol} \cdot \text{dm}^{-3}$ ). (3) Transfer The Above Solution Quantitatively To A 100  $\text{cm}^3$  Volumetric Feb 3th, 2024

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