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42 Chapter 2 Environmental Chemistry Section 2.2 Stoichiometry 43 We Already Know That There Are 44 Grams Per Mole Of CO" So We Do Not Need To Recalculate That. Two Moles Of Butane ($2 \text{ Mol} \times 58 \text{ G/mol} = 116 \text{ G}$) Yields 8 Moles Of Carbon Dioxide ($8 \text{ Mol} \times 44 \text{ G/mol} = 352 \text{ G CO}_2$) , So We Can Set Up The Following Proportion: 4th, 2024

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Theoretical yield = $5.95\text{g KClO}_4 \times \frac{1\text{mol}}{122.55\text{g}} \times 2 \times \frac{32.00\text{g}}{1\text{mol}} = 3.11\text{g O}_2$ %Yield = $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$ 34.
 $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$ Percent purity = 62.8%? pure FeCO_3 1th, 2024

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