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Chemistry Stoichiometry Problem Sheet 2 Answer Key Stoichiometry: Problem Sheet 1 - FREE Chemistry Materials Chemistry: Stoichiometry - Problem Sheet 1

Directions: Solve Each Of The Following Problems. Show Your Work, Including Proper Units, To Earn Full Credit. 1. Silver And Nitric Acid React According To The Following Balanced Equation:  $3 \text{Ag}(s) + 4 \text{HNO}_3(aq) \rightarrow 3 \text{AgNO}_3(aq) + 2 \text{H}_2\text{O}(l) + \text{NO} \dots$

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Cl 1mol 71 G Cl Cl X 546 G Cl 10) 292 G Ag 1 Mol Ag 108 G Ag 1 Mol Cu 1 Mol Ag 63.5 G Cu  
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1.  $\underline{\hspace{1cm}}$  CaCl<sub>2</sub> +  $\underline{\hspace{1cm}}$  AgNO<sub>3</sub> →  $\underline{\hspace{1cm}}$  Ca(NO<sub>3</sub>)<sub>2</sub> +  $\underline{\hspace{1cm}}$  AgCl  
How Many Grams Of Silver Chloride Are Produced When 45 G Of Calcium Chloride React With Excess Silver ...  
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Stoichiometry Problem Sheet 1 Answers  
Stoichiometry Volume Problems Worksheet Answers  
Stoichiometry: Mixed Problems (KEY)  
1) N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub>  
What Volume Of NH<sub>3</sub> At STP Is Produced If 25.0 Of N<sub>2</sub> Is Reacted With An Excess Of H<sub>2</sub>?  
3 3 3 2 3 2 2  
2 40.0L NH<sub>3</sub>  
1mol NH<sub>3</sub>  
22.4  
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Chemistry 11 Stoichiometry Review Worksheet Name: KEY  
33. KClO<sub>4</sub> → 2O<sub>2</sub> + KCl  
Actual yield = 1.45g  
Theoretical yield = 5.95g  
KClO<sub>4</sub> X 1mol / 122.55g X 2O<sub>2</sub> / 1KClO<sub>4</sub> X 32.00g / 1mol = 3.11g  
O<sub>2</sub> % Yield = 1.45g / 3.11g X 100% = 46.6%  
34. 4FeCO<sub>3</sub> + O<sub>2</sub> → 2Fe<sub>2</sub>O<sub>3</sub> + 4CO<sub>2</sub>  
Percent purity = 62.8%  
?g pure FeCO<sub>3</sub>  
12th, 2024

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