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63.28 G Of Phosphorus. Calculate The Empirical Formula. 9. Find The Molecular Formula Of Ethylene Glycol. The Molar Mass Is 62 G/mol And The Empirical Formula Is CH_3O . 10. A Compound Is Composed Of 7.20 G Of Carbon, 1.20 G Of Hydrogen, And 9.60 Grams Feb 2th, 2024
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Name: ___KEY___
33. $\text{KClO}_4 \rightarrow 2 \text{O}_2 + \text{KCl}$ Actual yield = 1.45g Theoretical yield = 5.95g
 $\text{KClO}_4 \times 1 \text{ mol} = 122.55 \text{ g}$
 $\text{O}_2 \times 2 = 32.00 \text{ g}$
 $1 \text{ mol} = 3.11 \text{ g}$
 O_2 % Yield = $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$
34. $4 \text{FeCO}_3 + \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 4 \text{CO}_2$
Percent purity = 62.8%? g pure FeCO_3 Jun 4th, 2024.

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Equation For The Reaction And Either The Masses Of Solid Reactants And Products
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14.2 Grams Of P₄O₁₀ From P? (Molar Mass P₄O₁₀ = 284) (A) 0.0500 Mole (B)
0.0625 Mole (C) 0.125 Mole Feb 3th, 2024

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