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3 Another Example A Solution With Two Volatile Components •Modified Raoult's Law • $VP_T = X_A P_A^* + X_B P_B^*$  Example •At 20.0 °C The Vapor Pressures Of Methanol ( $CH_3OH$ )and Ethanol ( $C_2H_5OH$ ) Are 95.0 And 45.0 mmHg Respectively. An Ideal Solution Contains 16.1 g Of Methanol Jan 8th, 2024

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Colligative Properties Are Properties That Depend Only Upon The Number Of Solute Atoms, Ions, Or Molecules In A Solution And Not On The Nature Of Those Atoms, Ions Or Molecules. Freezing Point Depression And Boiling Point Elevation Are Examples Of Colligative Properties. May 4th, 2024

## **Colligative Properties- Page 1 Lecture 4: Colligative ...**

B Raises Boiling Point Of Water From 100 °C To 100.64 °C •  $\Delta T_f$  Lowers Freezing Point Of Water From 0 °C To -2.32 °C •  $\pi$  Raises Osmotic Pressure From 1 Atm To 28.6 Atm While The Temperature Was Changing By A Degree May 6th, 2024

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Unpaid Property' 'Dekalb County Schools User Manual  
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4. The Normal Freezing Point Of Ethanol Was Obtained  
From The CRC Handbook Of Chemistry & Physics [3]  
And Found To Be  $-114.1^{\circ}\text{C}$  At 0.1 MPa (roughly  
Atmospheric Pressure). The Freezing Point Depression  
Of The Solution Of Methanol In Ethanol Was Then  
Calculated Using This Normal Freezing Point Of Ethanol  
And Equation (17). Apr 8th, 2024

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If You Add 30.0g Of Salt To 3.75kg Of Water, What Will Be The Change In The Boiling Point Of Your Salt Water? Assume The  $K_b$  Of Water Is  $0.510^\circ\text{C}/m$ .  $30.03 \times .51$  Assume You Have A  $3.60m$  Solution That Depressed The Freezing Point Of The Solution By  $0.8510^\circ\text{C}$ . What Is The Molal Freezing Poin Jun 2th, 2024

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Water (a) HCl C An Ionic Compound That Is Not Very  
Soluble In Water (b)  $\text{NaNO}_3$  Feb 2th, 2024

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