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Equation:  $E_{\text{Cell}} = (RT/nF) \ln K$  R Is Gas Law Constant, T Is Temperature, K Is Equilibrium Constant, F Is The Faraday, And N Is Number Of Equivalents Of Electricity Electrode Potential Is Sum Of Two Half-reactions. Cannot Determine Potential Of A Single Electrode, Measure Differences In Potential. File Size: 1MB Page Count: 107 Jan 4th, 2024.

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E Cell = E Ind - E Ref (+ E J) Nernst Equation  
 $E = E^{\circ} - \frac{RT}{nF} \ln \frac{[Red]}{[Ox]}$  Where R = Gas Constant T = Absolute Temperature N = Number Of Electrons In Reaction F = Faraday's Constant E = Potential E<sup>o</sup> = Standard Potential [Red] = Molar Concentration Of Reduced Form Of Speci

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