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$N\text{O} + n\text{e}^- \rightleftharpoons \text{NRR}$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G = \Delta G^0 + RT \ln \frac{a_{\text{N}} a_{\text{R}}}{a_{\text{H}^+} a_{\text{O}}}$ (2.1.38) Which Since $\Delta G = -nF E_{\text{rxn}}$ And $a_{\text{H}^+} = 1$ Becomes, $E = E^0 + \frac{RT}{nF} \ln \frac{a_{\text{N}} a_{\text{R}}}{a_{\text{O}}}$. (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can 2th, 2024

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4 1 Introduction To Electrochemical Cells Table 1.1 Dimensions Of Commercially Available Battery Sizes [1] . Battery Size Diameter (mm) Height (mm) N 12 30.2 AAA 10.5 44.5 AA 14.5 50.5 C 26.2 50 D 34.2 61.5 F 32.0 91.0 Flat Cells 1th, 2024

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