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$N\text{O} + ne \rightleftharpoons \text{NRR}$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G = \Delta G^0 + RT \ln \frac{a_{\text{N}} a_{\text{H}^+}}{a_{\text{NO}} a_{\text{H}_2}}$ (2.1.38) Which Since $\Delta G = -nF E_{\text{rxn}}$ And $\frac{a_{\text{H}^+}}{a_{\text{H}_2}} = \frac{a_{\text{H}^+}}{1} = 1$ Becomes, $E = E^0 + \frac{RT}{nF} \ln \frac{a_{\text{N}}}{a_{\text{NO}}}$. (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can 3th, 2024

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Elena Arroyo-de Dompablo, ZhaopingLiu,YingShirley Meng Liuzp@nimte.ac.cn (Z.L.) Shmeng@ucsd.edu (Y.S.M.) HIGHLIGHTS

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