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Excess Of Ethanol Was Used Because Its Lower Critical Point Reduces The Reaction Middle, Making This Reaction Technically Viable. The Estimates Of The Critical Temperature And Pressure Of The SODD And Ethanol Mixture Were Presented In Section 2.1 And The Values Are Shown In Table 2. At 553.15K, The Molar Ratios Of 10, 20 And 30 Lead 4th, 2024

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In Your Reaction, The Synthesis Of Banana Oil – Isopentyl Acetate, You Will Be Using Isopentyl Alcohol (3-methyl-1-butanol) And Acetic Acid (ethanoic Acid, See Figure 3).  $\text{H}_3\text{C}-\text{O}-\text{O}-\text{H} + \text{HO}(\text{CH}_2)\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2\text{OH} \rightarrow \text{H}_3\text{C}-\text{O}-\text{O}-\text{CH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2\text{OH} + \text{H}_2\text{O}$  Acetic Acid (ethanoic Acid) MW 60.1, D 1.05 Bp 118 °C Isopentyl Alcohol (3-methyl-1-butanol) MW 88.1, D 0.81 1th, 2024

### **22. The Fischer Esterification**

In This Laboratory Experiment, You Will Prepare Banana Oil (isopentyl Acetate) By The Fischer Esterification Of Acetic Acid And Isopentyl Alcohol (figure 3). The Reaction Will Be Catalyzed By The Addition Of Sulfuric Acid. Once The Reaction Is Complete, A Distillation Will Be Performed To ... 2th, 2024

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1. Esters Of High Volatility – Such As Methyl Formate Methyl Acetate. Ethyl Formate Having Lower Boiling Point Than Those Corresponding Alcohols And Therefore, Can Easily Be Removed From The Reaction Mixture. 2. Esters Of Medium Volatility – These Are Capable Of Removing Water ... 1th, 2024

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