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 $2 \rightarrow \text{CaSiO}_3$ B. $3\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow 2\text{Fe}_3\text{O}_4 + \text{CO}_2$ C. $\text{FeO} + \text{SiO}_2 \rightarrow \text{FeSiO}_3$ D. $\text{FeO} \rightarrow \text{Fe} + \frac{1}{2}\text{O}_2$ A. A B. D C. C And D D. A And D
 Answer: C
 Solution: In Metallurgy Of Iron, CaO Is Used As Flux Which Is Used To Remove The Impurities Of SiO_2 , $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$. Also Here Fe_2O_3 Is Reduced By CO To Fe_3O_4 Which Is Further Reduced To ... May 4th, 2024
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