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 Nitrogen And Hydrogen (see Section Chemical  
 Equilibrium Of The Thermodynamics Chapter). This  
 Reaction Follows The Equation: N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub> (1)  
 H<sub>0</sub> = 92 KJ Mol S<sub>0</sub> = 192 J Mol K Jan 1th, 2024  
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 Of fluxes Through Each Face Of The Volume \* D<sup>-b</sup> ← V  
 ← = F. K (t)+V ← R<sup>-</sup> ← (t) Dt • We Want To Solve For By  
 Approximating The Reaction And . flux Terms. Let's  
 Construct Low Order Approximations Physically. D B ←  
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