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The Percent Yield If 375.0 G Of Silver Bromide Was Obtained From The Reaction
Theoretical Yield = 407.9 G AgBr Percent Yield = $100 \times \frac{\text{Actual Yield}}{\text{Theoretical Yield}}$ 4th, 2024.
Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems) Stoichiometry
Worksheet #2 (mole-mass, Mass-mole Problems) 1. $\text{N}_2 + 2\text{O}_2 \rightarrow \text{N}_2\text{O}_4$ A. If
15.0g Of N_2O_4 Was Produced, How Many Moles Of O_2 Were Required? $15.0\text{g N}_2\text{O}_4 \times \frac{1\text{ Mol N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} = 0.163\text{ Mol N}_2\text{O}_4$
 $0.163\text{ Mol N}_2\text{O}_4 \times \frac{2\text{ Mol O}_2}{1\text{ Mol N}_2\text{O}_4} = 0.326\text{ Mol O}_2$ B. If
 4.0×10^{-3} Moles Of Oxygen Reacted, How Many Grams Of N_2 Were Needed?
 $4.0 \times 10^{-3}\text{ Mol O}_2 \times \frac{1\text{ Mol N}_2}{2\text{ Mol O}_2} \times 28\text{ g/mol} = 0.056\text{ g N}_2$... 4th, 2024

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