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4) -A Solution Is Created By Measuring 3.60×10^3 Moles Of NaOH And 5.95×10^{-4} Moles Of HCl Into A Container And Then Water Is Added Until The Final Volume Is 1.00 L. What Is The PH Of This Solution? Since There Is Both Acid And Base We Will Assume A 1 Mole Acid:1 Mole Base Ratio Of Neutralization. Jun 1th, 2024

PH And POH - Ms. Mogck's Classroom - Ms. Mogck's Classroom

KEY Chemistry: PH And POH Calculations Part 1: Fill In The Missing Information In The Table Below. PH $[H_3O^+]$ POH $[OH^-]$ ACID Or BASE? 3.78 $1.66 \times 10^{-4} M$ 10.22 $6.03 \times 10^{-11} M$ Acid 3.41 $3.89 \times 10^{-4} M$ 10.59 $2.57 \times 10^{-11} M$ Acid 8.81 $1.55 \times 10^{-9} M$ 5.19 $6.46 \times 10^{-6} M$ Base 8.69 $2.04 \times 10^{-9} M$ 5.31 $4.88 \times 10^{-6} M$ Base 8.46 $3.47 \times 10^{-9} M$ 5.54 $2.88 \times 10^{-6} M$ Base Jan 1th, 2024

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