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**Amanda Poh (n.) A·man·da·poh - SFU**

Creative Ideas For Existing Businesses. June – August 2015 Simon Fraser University  
Bachelor Of Arts In Design School Of Interactive Arts + Technology 2012 – 2018  
Inamoto & Co. Product Design Intern At I&Co., I Worked Closely With A Team Of 4 To  
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Hydrogen Ion Concentration, Shown As  $[H^+]$ , For Water =  $1.0 \times 10^{-7}M$  Hydroxide Ion Concentration, Shown As  $[OH^-]$ , For Water =  $1.0 \times 10^{-7}M$  Self-ionization Produces A Tiny Number Of Ions But Explains How Pure Water Can Behave As A Very Weak Electrolyte. D) Water Is The Usual Solvent For Acids And Bases; The Dissociation Of Acids Or Bases In An 3th, 2024

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4) -A Solution Is Created By Measuring  $3.60 \times 10^3$  Moles Of NaOH And  $5.95 \times 10^{-4}$  Moles Of HCl Into A Container And Then Water Is Added Until The Final Volume Is 1.00 L. What Is The PH Of This Solution? Since There Is Both Acid And Base We Will Assume A 1 Mole Acid:1 Mole Base Ratio Of Neutralization. 4th, 2024

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KEY Chemistry: PH And POH Calculations Part 1: Fill In The Missing Information In The Table Below. PH [ H 3 O 1+ ] POH [ OH1- ] ACID Or BASE? 3.78  $1.66 \times 10^{-4}$  M 10.22  $6.03 \times 10^{-11}$  M Acid 3.41  $3.89 \times 10^{-4}$  M 10.59  $2.57 \times 10^{-11}$  M Acid 8.81  $1.55 \times 10^{-9}$  M 5.19  $6.46 \times 10^{-6}$  M Base 8.69  $2.04 \times 10^{-9}$  M 5.31  $4.88 \times 10^{-6}$  M Base 8.46  $3.47 \times 10^{-9}$  M 5.54  $2.88 \times 10^{-6}$  M Base 4th, 2024

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