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 $7 + 6 \text{ KI} + 7 \text{ H}_2\text{SO}_4 \rightarrow \text{Cr}_2(\text{SO}_4)_3 + 4 \text{ K}_2\text{SO}_4 + 3 \text{ I}_2 + 7 \text{ H}_2\text{O}$
A) How Many Moles Of Potassium Dichromate ($\text{K}_2\text{Cr}_2\text{O}_7$) Are Required ... = 407.9 G AgBr This Is The Theoretical Yield Yields 22
B) Calculate The Percent Yield If 375.0 G Of Silver Bromide Was Obtained From The Reaction
Theoretical Yield = 407.9 G AgBr Percent Yield = $100 \times \frac{\text{Actual Yield}}{\text{Theoretical Yield}}$
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