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16. Find The Percent Of Ionization Of Propanoic Acid, CH 3CH 2COOH, Which Has A K A = $1.34 \times 10-5$. The Concentration Of The Propanoic Acid Is 0.100 M. 17. Find The PH Of A 0.075 M HCN Solution That Ionizes 0.00907%. 18. Find The PH Of Solutions With The Following H 3O + Concentrations. A) $1.15 \times 10-6 \text{ M}$ B Jan 2th, 2024

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